

# Chapter 7 Review Chemical Formulas And Compounds Section 2

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## **Chapter 7 Review Chemical Formulas**

CHAPTER 7 REVIEW Chemical Formulas and Chemical Compounds SECTION 1 SHORT ANSWER Answer the following questions in the space provided. 1. c In a Stock system name such as iron(III) sulfate, the Roman numeral tells us (a) how many atoms of Fe are in one formula unit. (b) how many sulfate ions can be attached to the iron atom. (c) the charge on each Fe ion.

## **7 Chemical Formulas and Chemical Compounds**

CHAPTER 7 REVIEW Chemical Formulas and Chemical Compounds MIXED

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REVIEW SHORT ANSWER Answer the following questions in the space provided. 1. Write formulas for the following compounds: a. copper(II) carbonate b. sodium sulfite c. ammonium phosphate d. tin(IV) sulfide e. nitrous acid 2. Write the Stock system names for the following compounds: a.  $\text{Mg}(\text{ClO}_4)_2$  b.  $\text{Fe}(\text{NO}_3)_2$

## **7 Chemical Formulas and Chemical Compounds**

CHAPTER 7 REVIEW Chemical Formulas and Chemical Compounds SECTION 1 SHORT ANSWER Answer the following questions in the space provided 1 c In a Stock system name such as iron(III) sulfate, the Roman numeral tells us (a) how many atoms of Fe are in one formula unit (b) how many sulfate ions can be attached to the iron atom (c)

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SHORT ANSWER Answer the following questions in the space provided. 1. c In a Stock system name such as iron(III) sulfate, the Roman numeral tells us (a) how many atoms of Fe are in one formula unit. (b) how many sulfate ions can be attached to the iron atom. (c) the charge on each Fe ion.

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CHAPTER 7 REVIEW Chemical Formulas  
and Chemical Compounds SECTION 2  
SHORT ANSWER Answer the following  
questions in the space provided. 1.  
Assign the oxidation number to the  
specified element in each of the  
following examples: a. S in  $\text{H}_2\text{SO}_3$  b. S in  
 $\text{MgSO}_4$  c. S in  $\text{K}_2\text{S}$  d. Cu in  $\text{Cu}_2\text{S}$  e. Cr in  
 $\text{Na}_2\text{CrO}_4$  f. N in  $\text{HNO}_3$  g. C in  $(\text{HCO}_3)$  h.  
N in  $(\text{NH}_4)$  2. a.

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Chapter 7 Significance of a Chemical Formula •A chemical formula indicates the relative number of atoms of each kind in a chemical compound. •For a molecular compound, the chemical formula reveals the number of atoms of each element contained in a single molecule of the compound. •example:

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octane —C<sub>8</sub>

## **Chemical Formulas and Chemical Chapter 7 Compounds**

A chemical formula includes the symbols of the elements in the compound and subscripts that indicate Chemistry Ch. 7 Review--Chemical Formulas and Chemical Compounds DRAFT 10th - 12th grade

## **Chemistry Ch. 7 Review--Chemical Formulas and Chemical ...**

CHAPTER 7 REVIEW Chemical Formulas and Chemical Compounds SECTION 2 SHORT ANSWER Answer the following questions in the space provided 1 Assign the oxidation number to the specified element in each of the following examples: 4 a S in H<sub>2</sub>SO<sub>3</sub> 3 6 b S in MgSO<sub>4</sub> 2 c S in K<sub>2</sub>S 1 d Cu in Cu<sub>2</sub>S 6 e Cr in Na<sub>2</sub>CrO<sub>4</sub> 5 f N in HNO<sub>3</sub> 4 g C in (HCO<sub>3</sub>)<sub>2</sub> [MOBI] Chapter 7 Modern Chemistry Review Answers

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## **Review Chemical Formulas ...**

CHAPTER 7 REVIEW. Chemical Formulas and Chemical Compounds. MIXED REVIEW. SHORT ANSWER Answer the following questions in the space provided. 1. Write formulas for the following compounds: a.copper(II) carbonate. b.sodium sulfite. c.ammonium phosphate.

## **CHAPTER 7 REVIEW**

CHAPTER 7 REVIEW Chemical Formulas and Chemical Compounds SECTION 1 SHORT ANSWER Answer the following questions in the space provided. 1. c In a Stock system name such as iron(III) sulfate, the Roman numeral tells us (a) how many atoms of Fe are in one formula unit. (b) how many sulfate ions can be attached to the iron atom.

## **Modern Chemistry Chapter 7 Review Chemical Formulas And ...**

chemical formulas.In this chapter,you will be introduced to some of the rules used to identify simple chemical



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compounds. Significance of a Chemical Formula Recall that a chemical formula indicates the relative number of atoms of each kind in a chemical compound. For a molecular compound, the

## **CHAPTER 7 Chemical Formulas and Chemical Compounds**

Using Chemical Formulas SECTION 4

Determining Chemical Formulas

CHAPTER 7 Untitled-2 206 6/28/2011

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Overview Section 1 describes the

naming of binary ionic and molecular

compounds. Section 2 describes the way

oxidation numbers are assigned and the

Stock system of naming compounds.

Section 3 describes how to calculate

### **1 Core Instruction**

CHAPTER 7 REVIEW Chemical Formulas

and Chemical Compounds SECTION 1

SHORT ANSWER Answer the following

questions in the space provided. 1. c In a

Stock system name such as iron (III)

sulfate, the Roman numeral tells us (a)

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how many atoms of Fe are in one formula unit. (b) how many sulfate ions can be attached to the iron atom.

## **Holt Modern Chemistry Chapter 7 Review Answer Key**

CHAPTER 7 REVIEW Chemical Formulas and Chemical Compounds SECTION 1  
SHORT ANSWER 1. In a Stock system name such as iron (III) sulfate, the Roman numeral tells us (a) (b) (c) (d) 2.

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Chemical Formulas and Chemical  
Compounds MIXED REVIEW SHORT  
ANSWER Answer the following questions  
in the space provided. 1. Write formulas  
for the following compounds:  $\text{CuCO}_3$  a.  
copper(II) carbonate  $\text{Na}_2\text{SO}_3$  b. sodium  
sulfite  $(\text{NH}_4)_3\text{PO}_4$  SnS<sub>2</sub> c. ammonium  
phosphate d. tin(IV) sulfide  $\text{HNO}_2$  e.  
nitrous acid 2.

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